

MARK SCHEME for the October/November 2013 series

0620 CHEMISTRY

0620/21

Paper 2 (Core Theory), maximum raw mark 80

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2013 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.

Page 2	Mark Scheme	Syllabus	Paper
	IGCSE – October/November 2013	0620	21

- 1 (a) (i) nitrogen [1]
- (ii) sulfur [1]
- (iii) iodine [1]
- (iv) helium [1]
- (v) nickel [1]
- (vi) iodine [1]
- (b) substance containing only 1 type of atom / substance which cannot be broken down further by chemical means [1]
- (c) Any 3 of: [3]
- conducts electricity / conducts heat / conducts
shiny / lustrous
ductile / can be drawn into wires
malleable / can be shaped
ALLOW: high boiling point / high melting point / solid at room temperature
ALLOW: rings when hit / sonorous
- [Total: 10]**
- 2 (a) (i) pair of bonding electrons [1]
8 electrons around chlorine and no additional electrons around hydrogen [1]
- (ii) covalent because has shared (pair of) electrons [1]
ALLOW: low melting point / low boiling point / it is a gas / doesn't conduct electricity / both non-metals
- (b) pH 2 [1]
- (c) (i) calcium chloride [1]
carbon dioxide [1]
water [1]
NOTE: do not allow formulae
- (ii) 2 [1]
calcium chloride [1]

Page 3	Mark Scheme	Syllabus	Paper
	IGCSE – October/November 2013	0620	21

(d) (i) values from 215 to 245 (s) [1]

(ii) 22 (cm³) [1]

(iii) Any 2 of: [2]

temperature / mass of magnesium / particle size of magnesium / surface area of magnesium

[Total: 13]

3 (a) 1 mark each correct answer [4]

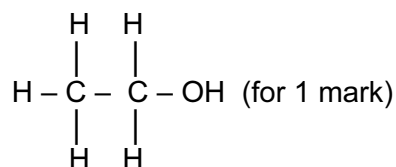
carbon / hydrogen

hydrogen (if carbon given for first marking point) / carbon (if hydrogen given for first marking point)

similar

functional

(b) (i)



(ii) carbon dioxide [1]

water [1]

(c) (i) COOH ringed [1]

(ii) 7 [1]

(iii) foodstuffs / drinks / cosmetics / water [1]

IGNORE: generalised answers e.g. kitchen / cleaning

[Total: 11]

Page 4	Mark Scheme	Syllabus	Paper
	IGCSE – October/November 2013	0620	21

4 (a) Any 4 of: [4]

both contain carbon atoms
 both have covalent bonding
 both are giant structures / lattices
 both contain rings / have hexagonal patterns / rings of 6 atoms
 in diamond, atoms arranged tetrahedrally
 in graphite, atoms arranged in layers
 flat rings in graphite
 bent rings in diamond
 all bonds same length in diamond
 graphite has some longer bonds / weaker bonds
 in diamond, each C atom joined to 4 others
 in graphite, each C atom joined to 3 others

(b) lime water; [1]
 turns milky / cloudy / white ppt [1]
 2nd mark dependent on correct reagent

(c) poisonous / kills you / toxic [1]
ALLOW: harmful / higher level answers referring to combining with haem
IGNORE: causes respiration problems / damages lungs

(d) oxygen removed from iron oxide [1]
ALLOW: oxidation number of iron decreases / iron gains electrons / CO becomes oxidised / oxygen adds to CO

(e) limestone [1]
 air [1]

[Total: 10]

5 (a) filter paper / chromatography paper [1]
 solvent / alcohol / other suitable solvent [1]
NOT: leaves / pigments in solvent

(b) X drawn on base line [1]

(c) chromatography [1]

(d) (i) 2nd box down ticked / aqueous nickel(II) sulfate [1]

(ii) nickel [1]

(iii) cathode [1]

Page 5	Mark Scheme	Syllabus	Paper
	IGCSE – October/November 2013	0620	21

- (e) protection from corrosion / make it less reactive / make it unreactive [1]
better appearance / more shiny [1]

- (f) (i) $6\text{H}_2\text{O}$ [1]

- (ii) reversible reaction / equilibrium reaction / reaction goes both ways /
reaction goes backwards as well (as forwards) [1]
IGNORE: reaction goes backwards / it is the reverse reaction

- (iii) add water (to white nickel(II) chloride) / hydrate (white nickel(II) chloride) [1]

[Total: 12]

- 6 (a) Any 4 of: [4]

in steam, molecules are far apart
in water, molecules are close together
in steam, molecules are moving very fast
in water, molecules are moving slowly / sliding over each other
in steam more randomness in arrangement of molecules
NOTE: molecules are further apart in steam (than in water) = 2 marks
NOTE: molecules move faster in steam (than in water) = 2 marks
NOTE: for molecules the word particles can be used
NOT: implication of particles 'apart' in liquids

- (b) (i) substance which dissolves another / it dissolves a solute / substance which dissolves a
solute / it dissolves something; [1]

- (ii) ethanol [1]
IGNORE: alcohol

- (c) endothermic [1]

- (d) 1st box ticked / aqueous ammonium chloride [1]

- (e) (i) LiOH on right [1]
2 on left (mark dependent on LiOH being correct) [1]

- (ii) 20 g [1]

[Total: 11]

- 7 (a) (i) copper [1]

- (ii) (copper is) better electrical conductor / iron is worse conductor [1]
IGNORE: copper is a good conductor

Page 6	Mark Scheme	Syllabus	Paper
	IGCSE – October/November 2013	0620	21

- (iii) does not conduct (electricity) [1]
- (iv) lead [1]
- (v) stronger / has more strength [1]
IGNORE: tougher / harder / less malleable
- (vi) lead [1]
- (b) (i) zinc [1]
- (ii) (zinc) hydroxide [1]
ALLOW: error carried forward from wrong metal in part (b)(i)
- (c) C,B,D,A [1]
- (d) CuCl_2 [1]
ALLOW: Cl_2Cu
- (e) positive electrode: chlorine [1]
negative electrode: copper [1]
ALLOW: 1 mark for chlorine and copper reversed
- (f) chlorine / Cl_2 [1]

[Total: 13]